

## Atomic Structure Chapter 4

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Chapters 4: Atomic Structure - ANNE SCHMIDT CHEMISTRY

Unformatted text preview: Chapter 4 "Atomic Structure" Pearson Chemistry Book High School 1 OBJECTIVES: 1. Defining the atom. 2. Recognize the change of the atomic structure through history. 3. Differentiate between the properties of subatomic particles. 4. Describe the characteristics of the atom. 2 I. Defining the Atom Warm up ( 1 min ) What is an atom? 3 Living things are made up of Cells ...

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4.05 Atomic Number, Mass Number, Isotopes and Isobars 4.05 Atomic Number, Mass number, Isotopes and Isobars. Chapter 4 Structure of Atoms - Test. 05 The Fundamental Unit of Life 5 Topics | 6 Quizzes 5.01 Introduction to cell 5.01 Introduction to Cell. 10.03 Plasma Membrane 5.02 Plasma Membrane ...

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1. All elements are composed of tiny invisible particles (atoms) 2. All atoms of the same element are identical. 3. Atoms of different elements can physically combine together or chemically combine in simple whole number ratios to form compounds. 4. Chemical reactions occur when atoms are separated, joined, or rearranged.

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Dalton's Atomic theory. 4 part theory about atoms. 1) All elements are composed of \_\_\_\_\_. atoms. Atoms can combine both \_\_\_\_\_ and \_\_\_\_\_. physically & chemically. 2) atoms in a given element are \_\_\_\_\_, but different from those in other elements. identical. 3) atoms are neither \_\_\_\_\_ nor \_\_\_\_\_.

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Chemistry- Chapter 4 Atomic Structure. Democritus's Idea. Rutherford's explanation of atomic struc.... JJ Thomson's. Atoms, element. Atoms are indivisible and indestructible. Nucleus is small, dense, and positively charged (is the explin.... Whose Experiment proved atoms are neutral, electrons have very....

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The nucleus has two fewer protons, so the atomic number decreases by two. The nucleus also has two fewer neutrons, so the mass number decreases by four. A new element is formed.

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Chapter 4 Atomic Structure Answers

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Chapter 4 Atomic Structure Workbook Answers

The central nucleus of the atom contains the neutrons and positively charged protons. The nucleus is surrounded by as many orbital electrons (each with a charge of  $-1$ ) as there are. protons (each with a charge of  $+1$ ) in the nucleus. The three sub-atomic particles differ from each other in two important ways.

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